beyond doubt that the strong bands are associated only with the mixed-valence ions. However, the electronic spectrum of the (μ-pyrazine)-*cis*-(dinitrogen)nonaamminediosmium 5+ ion is dramatically different, approximating a summation of spectral properties of a tetraammine Os(II) coordinated by dinitrogen and pyrazinium ion and a pentaammine Os(III) pyrazinium center.[6,7] Further, the dinitrogen stretching frequency (2075 cm^{-1}) is appropriate for Os(II), and a band with comparable extinction occurs at 1600 cm^{-1}. This band has been previously associated with a ring mode of the pyrazine, which is normally IR inactive for D_{2h} symmetry but becomes allowed when the center of symmetry has been lost, as in localized mixed-valence systems.[18] This 1600-cm^{-1} peak is quite weak in $[(NH_3)_4ClOs(pyz)Os(NH_3)_5]^{4+}$ and is virtually absent in the symmetrical $([(NH_3)_5Os]_2pyz)^{5+}$ ion, in harmony with their delocalized description. A detailed spectra analysis is now in progress[19] to further our understanding of the electronic structures in relation to both a molecular orbital description and the effects of spin-orbital coupling.

Acknowledgment. Support of this work by National Science Foundation Grant CHE79-08633 and National Institutes of Health Grant GM13638-17 is gratefully acknowledged. P.A.L. also acknowledge the receipt of a CSIRO Postdoctoral Fellowship.

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(18) Callahan, R. W.; Keene, F. R.; Meyer, T. J.; Salmon, D. J. *J. Am. Chem. Soc.* **1977**, *99*, 1064–1073.

(19) Lay, P. A.; Magnuson, R. H.; Taube, H., work in progress.

(20) It was necessary to use the Kirsten–Dumas method to obtain satisfactory analytical data for nitrogen. (Lay, P. A.; Sargeson, A. M.; Skelton, B. W.; White, A. H. *J. Am. Chem. Soc.* **1982**, *104*, 6161–6164.) All microanalyses were performed by the Stanford University Microanalytical Laboratory.

(21) Beattie, J. K.; Hush, N. S.; Taylor, P. R. *Inorg. Chem.* **1976**, *15*, 992–993.

Additions and Corrections

Polymerized Phosphatidylcholine Vesicles. Synthesis and Characterization [*J. Am. Chem. Soc.* **1982**, *104*, 791]. STEVEN L. REGEN,* ALOK SINGH, GÜNTHER OEHME, and MANINDER SINGH.

Page 793: Figure 5 is in error. The turbidity of both polymerized and nonpolymerized vesicles of **5** decreases approximately 30% on going from 0% to 25% ethanol. When 45% ethanol in water is used, nonpolymerized vesicles are completely destroyed (no apparent absorbance at 400 nm) and the polymerized dispersion exhibits a turbidity that is 20% higher than that found in pure water. We presume that this increase in turbidity is due to vesicle aggregation. In addition, phosphorus analysis now reveals a small but detectable extractability ($CHCl_3$) of lipid monomer, oligomer, and/or polymer from aqueous polymerized dispersions of **5**.

Cyclohexaamylose Complexation with Organic Solvent Molecules [*J. Am. Chem. Soc.* **1982**, *104*, 6283–6288]. ROBERT I. GELB, LOWELL M. SCHWARTZ,* MICHAEL RADEOS, ROBERT B. EDMONDS, and DANIEL A. LAUFER.

In our recent paper[1] we regretfully failed to mention a previous publication by Matsui and Mochida[2] which reports studies of some of the same chemical systems. Both papers determine aqueous complexation constants of cyclohexaamylose with ethanol, 2-propanol, 2-methyl-2-propanol, and cyclohexanol. Although the two studies employed entirely different experimental techniques, the sets of equilibrium constant values are in essential agreement.

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(1) Gelb, R. I.; Schwartz, L. M.; Radeos, M.; Edmonds, R. B.; Laufer, D. A. *J. Am. Chem. Soc.* **1982**, *104*, 6283–6288.

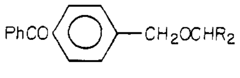
(2) Matsui, Y.; Mochida, K. *Bull. Chem. Soc. Jpn.* **1979**, *52*, 2808–2814.

Reactive Intermediates. Volume 2 [*J. Am. Chem. Soc.* **1982**, *104*, 7394].

First column, second line from the bottom, the sentence starting here should read as follows: An idea of the activity in these various fields is given by the fact that the carbene, free-radical, and silylene chapters are by far the longest, whereas that on arynes is less than a fifth as long as the carbene chapter.

Surface Photochemistry: Deviation of the Course of Reaction in Benzoin Ether Photolysis by Adsorption on Silica Gel [*J. Am. Chem. Soc.* **1982**, *104*, 6824–6825]. PAUL DE MAYO,* AKIRA NAKAMURA, PETER W. K. TSANG, and S. KING WONG.*

Page 6825, Scheme I: Structure **6** should be



6

Definitive Evidence for Cycloheptatetraene from Dehydrobromination of Bromocycloheptatrienes [*J. Am. Chem. Soc.* **1982**, *104*, 7329–7330]. JAMES W. HARRIS and W. M. JONES.*

Page 7329: In footnote 3 the date should be 1970 not 1980.

Iron(III)–Porphyrin π-Cation Radical Complexes. Molecular Structures and Magnetic Properties [*J. Am. Chem. Soc.* **1982**, *104*, 6793]. GEORGES BUISSON, ALAIN DERONZIER, EMILE DUEE, PIERRE GANS, JEAN-CLAUDE MARCHON,* and JEAN-RENE REGNARD.

Page 6795, second column, the second paragraph should read: Pertinent bond lengths in the coordination unit of **3** are as follows: $Fe–N_1$, 2.04 (1); $Fe–N_2$, 2.05 (2); Fe–O, 2.13 (1) Å. The equatorial bond length values are consistent with one-electron occupancy of the $d_{x^2-y^2}$ orbital, and therefore they confirm the high-spin Fe(III) state, ...

Kinetics, Thermodynamics, and Mechanism of the Radical Chain Process for Ligand Substitution of Metal Carbonyls [*J. Am. Chem. Soc.* **1983**, *105*, 61–73]. J. W. HERSHBERGER, R. J. KLINGLER, and J. K. KOCHI.*

Page 63: Entry 18 of Table II for $(MeC_5H_4)Mn(CO)_3$ should be corrected to—$E_p^{ox} = 1.15$ V and $(E_p^{ox} + E_p^{red})/2 = 1.10$ V.

Organic Electronic Spectral Data. Volume XVII [*J. Am. Chem. Soc.* **1982**, *104*, 7673].

The third editor of this reviewed book should be H. Feuer.